

H. Tuthill, B. Todd, J. Oldfield:

EXPERIENCE WITH COPPER ALLOY TUBING, WATERBOXES AND PIPING IN MSF DESALINATION PLANTS,

Proceedings IDA World Congress on Desalination and Water Reuse, Volume I, Sessions 1 to 3, October 6-9, 1997, Madrid, Spain,

ABSTRACT

The first generation MSF desalination plants were built with copper alloy tubing and either bare or coated carbon steel chambers, water boxes and piping. The intended design life was 20 years. It soon became apparent that inleakage of air into the chambers allowed more corrosion of steel than could be tolerated. Steel in these applications began to be replaced with clad or solid copper alloy, and in some plants with clad or solid stainless steel. Currently many units use copper alloy throughout, although titanium tubing is used in some heat reject sections where sulphides or special considerations are overriding considerations.

Today the industry is building ever larger plants and extending design life to 40 years or more. This paper reviews the data and experience with copper alloy tubing, water boxes, piping and chamber walls in MSF desalination plants. These data summarise what has been learned through the past 30 years usage of copper alloys which have become the basic material of construction for both land based and marine MSF desalination plants. This information should provide the current generation of designers engineers with the data and performance history needed to obtain ever greater reliability from these key materials in the future than in the past.

NATURE OF COPPER NICKEL ALLOYS

The copper nickel alloys are solid solutions of copper and nickel with an iron addition which enhances their resistance to erosion and turbulence. The compositions of the three most widely used copper nickel alloys are given in Table 1. The C70600 alloy is used for piping, tubesheets, waterboxes, and flash chambers as well as tubing. In flash chambers and water boxes, it is used as a lining, as clad or solid alloy depending upon economics and local preferences. The C71500 alloy is used as tubesheets and tubing and in slightly modified form as the welding filler metal. The C76140 alloy is used as tubing. Mechanical properties are given in Table 2.

Table 1: Nominal Composition in % of Copper Nickel Alloys

Alloy	Nickel	Iron	Mn	Cu
C70600*	9.0-11	.0 1.0-1.8	1.0 Max	Remainder
C71500*	29.0-33.0	0.4-1.0	1.0 Max	Remainder
C71640	29.0-32.0	1.7-2.3	1.5-2.5	Remainder

*For welding the following Max % are allowed Zn 0.50, Pb 0.02, P 0.02, S 0.02, and C 0.05.

Table 2: Mechanical Properties of Copper Nickel Alloys

		TS		YS	
		MPa	Ksi	MPa	Ksi
C70600	annealed	275	40	105	15
C70600	light drawn (tube)	310	45	240	35
C71500	annealed	310	52	110	18
C71640	annealed	435	63	170	25
C71640	drawn, stress relieved	560	81	400	58

Although their good thermal conductivity gives copper nickel a decided advantage over ferrous alloys and titanium, the principal characteristic that has made copper nickel, the alloys of choice for marine heat exchanger applications, is their outstanding resistance to biofouling. No other competitive material can provide such good heat transfer with so little use of chlorine or other biocides.

Lewis demonstrated this characteristic of copper nickel by replicating typical condenser cleaning cycles needed to maintain heat transfer capability in clean seawater with normal biofoulers present.(1) He found that C70600 needed to be cleaned of microfouling in about 110 days in order to maintain heat transfer capability, whereas a tube material with no inherent resistance to biofouling such as titanium would have to be cleaned 12 times in the same 110 day period to maintain equivalent heat transfer capability. Refer to Figure 1. This means, that if the use of chlorine were banned, or so reduced as to be ineffective, as many have advocated for environmental considerations, alloys without inherent resistance to biofouling would be at a serious disadvantage in the heat reject section. Lewis work also shows how dependent upon an adequate supply of chlorine are the alternative stainless steel and titanium tube materials.

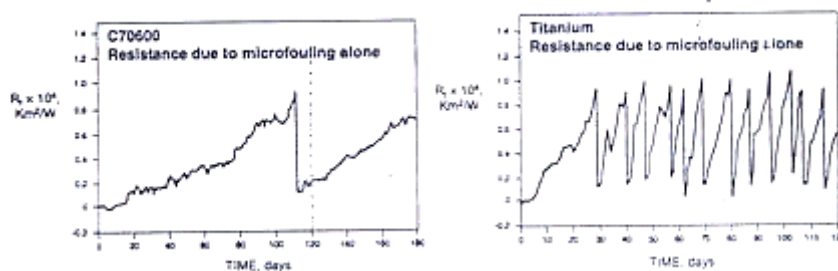


FIGURE 1: Time between cleanings to restore heat transfer due to microfouling in clean seawater (1).

WELDING

Although essentially solid solution alloys some precipitation of iron occurs in the C70600 alloy. The heat affected zone remains ductile and as strong as the base plate. Corrosion re-

sistance is not significantly affected by iron precipitation when it occurs. These inherent characteristics allow good quality welds to be made by all of the normal welding methods with only routine welder training. The 30 % Ni filler metal is used to weld both the 30% nickel and the 10% nickel base metal in order to avoid shop and field mixup problems that could lead to welding the 30% nickel alloy with a 10% nickel filler metal. Mechanically such a weld would be satisfactory, however such a weld which is essentially a heterogeneous cast structure, would corrode very rapidly galvanically due to using a less noble filler metal.

CORROSION BEHAVIOR

The corrosion resistance of copper nickel alloys depends upon formation of a protective film. Film formation is referred to sometimes as "passivation". Film formation is affected by pH, time, aeration, velocity, temperature, pollution and other factors. There is generally an inner cuprous oxide film, Cu_2O , and an outer cupric oxide, CuO , film.(2,3,4) Although cuprous and cupric oxides are the principal components of the films, the lattice usually includes other metallic ions, iron, nickel, aluminium, calcium, sometimes silicon, and sometimes other species. Principal anions include chlorides, hydroxides and carbonates, bicarbonates and oxides. There is no fixed composition.

The inner film is normally reddish. The outer film may be greenish, brown or yellow-brown. Despite variations in composition, these oxide type films offer a high degree of protection, when formed under favourable conditions. In polluted waters, and in well water from shore side wells, the film is often a sulphide, rather than an oxide, type film. Sulphide films are readily distinguished by their black colour. Sulphide films, while protective, are thicker, more porous and more easily damaged than oxide films.

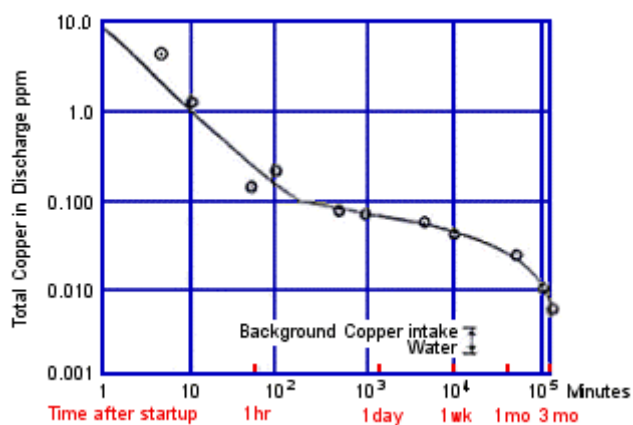


FIGURE 2 - Formation rate of corrosion product film on Alloy C70600 in 60F(16) seawater.

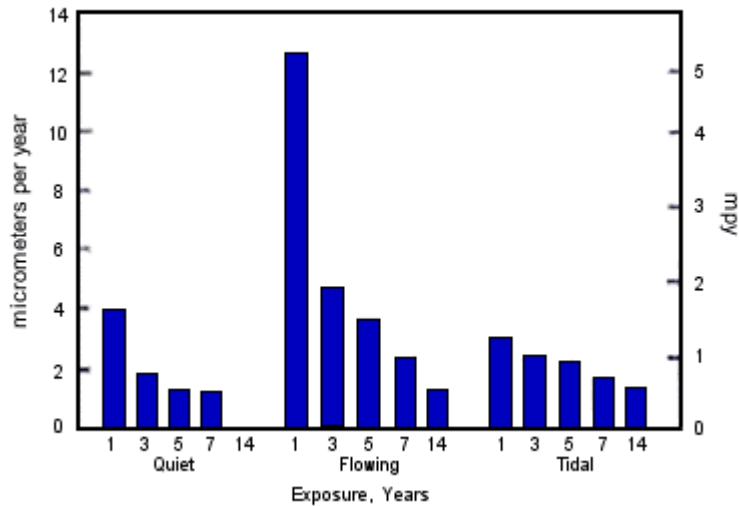


FIGURE 3 - Corrosion rates for Alloy C70600 for long term seawater exposures.

Time is a major factor in film formation and also in the degree of protection the film affords. Figure 2 shows the rate of film formation on C70600 in 60F(16C) seawater as measured by copper in the effluent.(5) In 10 minutes copper in the effluent decreased tenfold; in 2 hours a hundredfold, but it was 3 months before the film became "mature"; mature in the sense that copper in the effluent was only slightly more than copper in the intake water. Efirid and Anderson's data in Figure 3 shows that corrosion rate, as measured by weight loss, continues to decrease in quiet and flowing seawater well beyond 3 months up to, at least, 7 years.(6) The fact that the film becomes more protective with time accounts for the considerable variation in corrosion rates and durability of copper base alloys as reported by different investigators in the published literature. Data obtained in short term tests, before a fully mature film has formed, can be expected to overstate the corrosion rate and to understate the durability of copper alloys.

Temperature has a major influence on the rate of film formation. At 60F(16C), Figure 2, it took slightly over one hour for the film to cover the surface and 3 months before a reasonably mature film reduced copper in the discharge close to the level of copper in the intake. At higher temperatures the film forms and matures faster. At lower temperatures, the film forms and matures more slowly. At 35F(2C) it took a week before a visible film completely covered specimens of C70600 in some work the author undertook. This corresponds roughly with the slightly over one hour it took for full coverage to occur at 60F(16C), Figure 2. Although the film forms more slowly at low temperatures, it does form and become protective, even in arctic and antarctic waters, as demonstrated by the good performance of copper alloy tubed desalination plants at Prudhoe Bay, Alaska, and McMurdo Sound, Antarctica. At 80F(27C), a common inlet temperature for MidEast desalination plants, rapid film formation and good protection can be expected within a few hours.

Oxygen is very important in film formation on copper alloys. Oxygen reduction is the primary cathodic reaction supporting corrosion, and film formation, of copper alloys in seawater. In the complete absence of oxygen, there would be no film formation and no corrosion of copper alloys in most waters. The feed from the heat reject section to the heat recovery section of many MSF desalination plants is deaerated, hence corrosion in the heat recovery section was not originally expected. However, the flash chambers of the heat recovery section are under vacuum and air is sucked in. In actual practice, there seems always to be sufficient oxygen present due to the residual oxygen content of heat recovery section feed and to in-leakage of air, to support some corrosion of copper alloys as well as substantial corrosion of carbon steel.

Another major factor influencing film formation is pH. Ross and Ross and Anderson (7,8) studied film formation on C70600, C71500 and C68700 in seawater and found no film formation below pH 6 even though there was adequate oxygen, 5.5-7.5 mg/l, present. The un-filmed corrosion rates were high, of the order of 35 mpy(0.89mm/yr.). At higher pH's, where some oxygen was present, corrosion rates were low and normal.

Velocity has several effects. At velocities below 3fps(0.9mps), sediment settles out readily in horizontal runs of pipe and tubing. Unless removed periodically, under-sediment corrosion, as evidenced by a general pitting attack, may occur. Slime layers also build up at low velocities providing a good habitat for biofouling organisms to attach to the slime layer and grow. At normal heat exchanger tube velocities, much of the biofouling is swept away because it is so lightly attached to the slime layer, which is one factor tending to keep the tubes clean and heat transfer capability in an acceptable range.

At normal design velocities of 5-8fps(1.5-2.4mps), turbulence around debris lodged in copper alloy tubes, partially blocking flow, has led to pinhole type failures of copper alloy tubing downstream of such lodgements. Efirid studied the critical velocity at which the film is stripped away from copper alloys and the tube-wall-to-flowing seawater shear stresses that exist at the inlet end, 5 diameters down the tube, and around obstructions.(9) The critical velocity at which the protective film is stripped away by flow for the tube alloys is shown in Table 3. The tube-wall-to-flowing-seawater shear stresses are shown in Table 4. Efirid developed these data on specimens placed parallel to the flow in the centre of a one inch square cross section tube with clean sea water flowing on both sides of the specimens. The critical velocity for copper nickel is greater than the inlet end shear stress which confirms the better inlet end resistance of copper nickel as compared to other copper base alloy. The extreme turbulence that occurs around lodgements explain why pinhole type penetrations occur downstream of debris lodgements and underline the need for good screens and good screen maintenance.

Table 3. Critical velocity and shear stress for copper-base alloys in seawater.(9)

Alloy		Critical Velocity		Temp °C	Critical Shear Stress, Mpa
		m/s	ft/s		
C12200	(copper)	1.3	4.4	17	9.6
C68700	(aluminum brass)	2.2	7.3	12	19.2
C70600	(90/10 Cu/Ni)	4.5	14.7	27	43.1
C71500	(70/30 Cu/Ni)	4.1	13.5	12	47.9
C72200	(Cu-16Ni-0.5Cr)	12.0	39.4	27	296.9

Table 4. Shear stresses developed in condenser tubes in seawater.(9)

Design Velocity		Tube Wall Shear Stress, Mpa		
m/s	ft/s	Inlet	>5 L/D Down Tube	Around a 90% Blockage
2.0	6.5	18.4	10.8	264.6
2.5	8.1	28.0	<18.0	
2.75	9.0	33.0	20.0	
3.0	9.8	37.4	22.0	

SULFIDES

There have been a number of studies on the effect pollution and sulphides have on corrosion behaviours. There is considerable variation in these studies on the effect of sulphides and pollution. Syrett has published one of the better and most thorough studies on the effect of sulphides.(10)

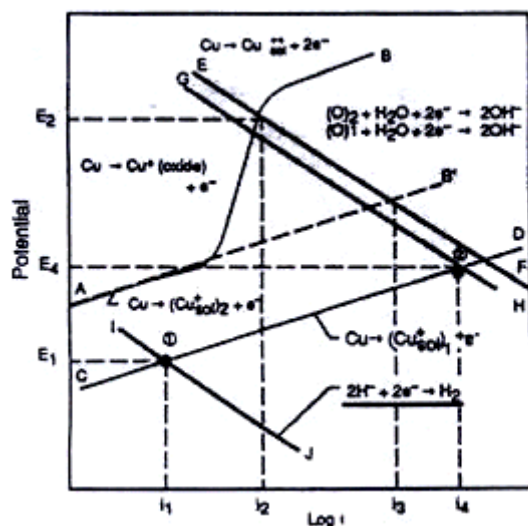


FIGURE 4 - Influence of sulphide and oxygen on the corrosion current in a copper-nickel alloy exposed to flowing seawater.

Figure 4 from Syrett's paper presents a large amount of useful information and is worth reviewing in some detail. The upper line AB represents the anodic reaction in aerated water and traces normal film formation. The initial section of line AB represents corrosion of un-

filmed metal. The next, nearly vertical section, depicts the initial film forming stage where cuprous oxide is forming. The intersection with the oxygen reduction line, identifies the steady state corrosion current i_2 at the normal potential, E_2 . The steady state corrosion current i_2 corresponds to the steady state corrosion rate which is normally less than 25 microns per year for this copper alloy.

Line CD represents the anodic reaction in deaerated waters where sulphide is present. The cathodic reaction is hydrogen reduction line IJ; the corrosion potential is E_1 ; and the steady state corrosion current is i_1 , point 1. The film that forms in fully deaerated, sulphide bearing waters is black, more porous, less adherent, and more easily damaged than the oxide type films previously described. The corrosion current is much lower, than in aerated waters, which would mean lower corrosion rates in sulphide infested waters, were it not for more porous, more easily damaged nature of the sulphide film. It should be noted that in deaerated waters in the absence of sulphides, hydrogen reduction does not occur, since copper will not displace hydrogen from water. It is only in deaerated sulphide bearing waters that the potential is depressed sufficiently for hydrogen reduction to occur.

When air is present in waters containing hydrogen sulphide, as can be the case with air in-leakage into the flash chambers which operate under a vacuum, the corrosion greatly accelerated. The intersection of the anodic line CD with the oxygen reduction line GH, point 2, identifies the condition where air and sulphides coexist in seawater. The corrosion current at point 3 is two orders of magnitude greater than i_2 , the corrosion current in normal aerated sea water, and, at least, three orders of magnitude greater than the i_1 , the corrosion current in deaerated sulphide bearing waters. Air and hydrogen sulphide do not coexist at equilibrium, as oxygen reacts with hydrogen sulphide to produce sulphur and water. However, there is abundant evidence that oxygen and hydrogen sulphide do coexist under transitory, non-equilibrium conditions. Corrosion can be quite severe when these conditions exist. For example, May and Weldon found that 3 ppm of H_2S continuously added to clean aerated seawater increased the corrosion rate of these copper alloys 4 to 10 times. (11)

AMMONIA

Ammonia is sometimes encountered in the seawater feed to desalination plants. Ammonia may come from agricultural run off or from industrial plant discharges near desalination plant intakes. Ammonia affects different copper alloys differently. In the presence of air and ammonia, aluminium brass is subject to stress corrosion cracking. Aluminium bronze is more resistant and copper nickels are highly resistant to ammonia stress corrosion cracking as Thompson has shown, Table 5 (12). Ammonia also tends to increase the general corrosion rates of copper alloys. Quantitative data for the increase in corrosion rate due to ammonia contamination is not yet available. However, Figure 5 from Michels et al, shows the copper

nickel alloys to be three orders of magnitude more resistant than aluminium brass, to general corrosion in spray and fog tests with 1000-2000 ppm ammonia present, a condition which can occur in the gas extraction zone of turbine condensers using hydroazine-type compounds for boiler water treatment.(13)

Table 5. Ammoniacal stress corrosion resistance of brasses, aluminum bronzes, and copper-nickel alloys. (12)

	Time to 50% relaxation (h)
Alloy	relaxation (h)
C44300	0.30
C28000	0.35
C46500	0.50
C26000	0.51
C68700	0.60
C60600	4.08
C61400	5.94
C70600	234
Cl4200 PDO copper	312
C71500	2000

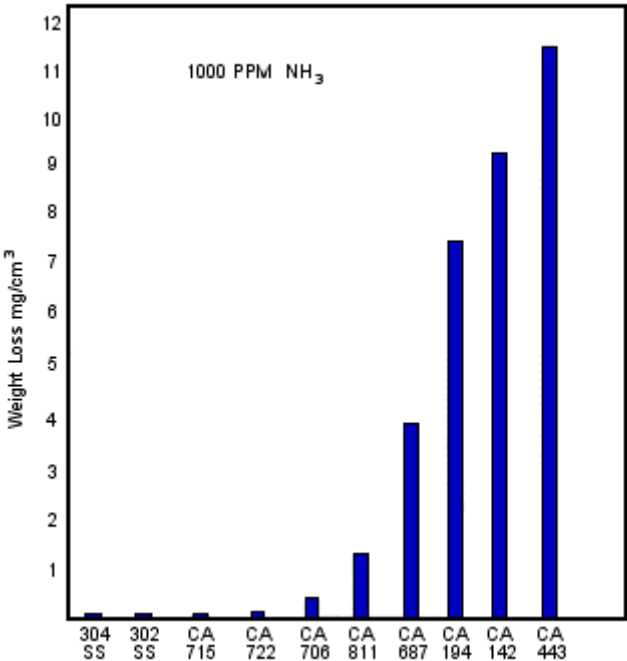


FIGURE 5 - Weight losses of various alloys after testing for 10 days in 1000 ppm NH3 contaminated condensate at 38C (100F). (13)

CHLORINE

Chlorination practices vary considerably from site to site, as both the need for chlorination and local governmental regulations affect the amount actually used. The effect of chlorine residuals in the 0.5 to 3.0 ppm range on three tubing alloys is shown in Table 6.(14). The corrosion rate of C68700 is increased substantially by residual chlorine in the 0.5-3.0 ppm, C70600 is slightly increased and C7 1500 hardly at all. Francis found the principal effect the ranges normally used was to increase the susceptibility of copper nickel to impingement attacks.(15) Overdosing beyond 3ppm, which unfortunately occurs from time to time in a number of desalination plants, can be detrimental to performance of copper alloy tubing. Tubing in the heat reject stage is more likely to be affected by overdosing. Tubing in the heat recovery stage downstream of the heat reject stage, is less affected by overdosing, as the residual chlorine is consumed rapidly. Chlorine reacts rapidly with organic material and oxidizable species in seawater.

Table 6. The effects of chlorination on copper tubing alloys - mpy (mm/y).(14)

Chlorine Residual	C68700	C70600	C71500
0.0 ppm	<1.0 (<0.03)	<1.0 (<0.03)	<1.0 (<0.03)
0.5 ppm continuous	2.9 (0.073)	1.6 (0.04)	0.4 (0.01)
0.5 ppm 2hr on 2 hr off	3.1 (0.08)	0.7 (0.02)	-
1.0 ppm 1 hr on, 5 hr off	0.9 (0.023)	3.2 (0.08)	0.5 (0.013)
1.5 ppm 2hr on, 2hr off	0.9 (0.023)	0.5 (0.013)	0.1 (0.003)
3.0 ppm 2 hr on, 2 hr off	2.4 (0.06)	0.3 (0.008)	0.2 (0.005)

SAND

Sand is common in seawater from open channels in shallow waters. High sand loadings can be detrimental to copper alloy tubing. Sato and Nagata have developed some useful guidelines on the effect of sand.(16). High sand loadings can damage the protective film, exposing bare metal to high unflmed corrosion rates. For any given sand loading, damage increases with the size and sharpness of the sand present.

The general order of copper alloy resistance to sand follows.

- 1) C68700 - least resistance
- 2) C70600 - more resistant
- 3) C71500 - more resistant
- 4) C7 1640 - outstanding resistance

Sand is not always detrimental. Fine beach sands which are the most likely to be encountered are less damaging than larger angular particles encountered in some estuaries. Rough guidelines follow.

1) Up to 200 ppm of sand - Unlikely to be a problem

2) 200-1000 ppm of sand - Increasing potential for damage. Good performance from the less resistant alloys has been reported after a mature film forms under favourable conditions.

3) >1000 ppm of sand - resistance of the first three alloys listed above is often inadequate. C7 1640 is usually resistant.

Intermittent exposure to even the higher loadings is not necessarily damaging, particularly when fully mature films have formed. Spikes are not as damaging as steady state loadings.

BRACKISH AND FRESH WATERS

Corrosion behaviour in fresh, brackish and higher salinity waters is quite similar to performance in seawater. Experience indicates that velocity and turbulence damage may be a little less in brackish waters than in seawater, but there have been few studies made of velocity effects in fresh or low salinity brackish waters. Copper alloys seem to perform much the same in low as well as high chloride waters. It is the other factors discussed above, rather than chloride ion concentration, that appear to control performance.

DESALINATION ENVIRONMENTS

The principal desalination environments and their characteristics are shown in Table 7.

Behaviour in the two heat reject environments where, clean aerated or deaerated, sulphide seawater environments are encountered, 1 and 2 in Table 7 has been described above. One additional comment is needed to cover performance in the heat reject sections. This comment also applies to heat recovery and brine heater tubing. Earlier practice in acid dosed plants, involved dropping the pH to 3-4 with sulphuric acid additions for 24 hours or so to clean the tubes. This was a very effective cleaning operation, so effective in fact, that the protective passive film was removed to virgin metal, and had to be reformed after each acid cleaning. Metal loss for copper alloy tubing during film formation is of the order of 0.5 to 1 mil (0.025 mm). While no tube failures have been reported due to such acid cleaning, estimated corrosion rates from wall thickness loss as measured in eddy current testing have been reported, and give unrealistically high corrosion rate information. In such acid dosed plants, it would be more realistic to report the metal loss per acid cleaning, rather than as an elapsed time corrosion rate.

Behaviour of copper alloy tubing in the heat recovery and brine heater sections was studied extensively in the Office of Saline Water desalination test facility at Dow Chemical Company's Freeport, Texas plant. (18) Figure 6 shows the flow diagram for this plant. Very careful

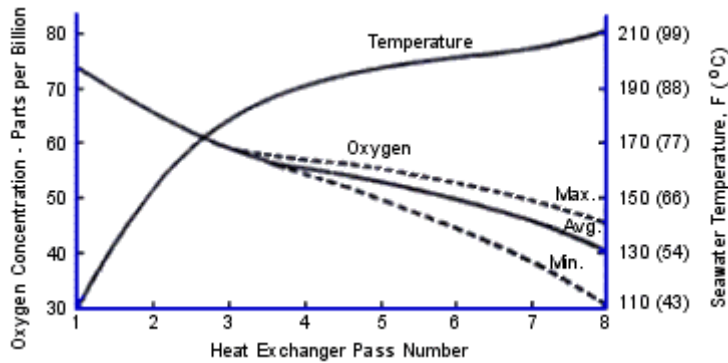


FIGURE 7 - Oxygen concentration and temperature of seawater in the heat recovery exchanger.(18)

Instantaneous corrosion rates for these conditions over 30 months of operation are shown in Figure 8. The corrosion rate for C68700 alloy decreased from its high initial rate and levelled off at 3 mpy (0.075 mmpy) at 20 months. The corrosion rates for C70600, C71500 and C61300 were lower and still decreasing after 30 months. These data indicate that longer service life could be expected from C70600, C7 1500 and C6 1300, than from C68700, in the heat recovery section of the plant. These data support the widespread preference for C70600 alloy tubing in the heat recovery section.

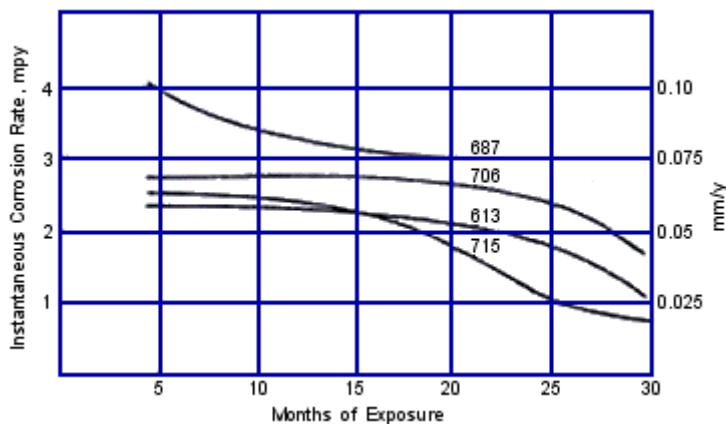


FIGURE 8 - Instantaneous corrosion rates versus exposure time of five copper alloys in heat recovery section of test plant.(18)

Corrosion rates for C68700, C70600 and C7 1500 in the brine heater section, shown in Figure 9 were low, about 1 mpy (0.025 mmpy) after 54 months. Oxygen was measured and reported to be zero in the brine heater. However, the fact that there was some corrosion indicates enough oxygen must have been present to support the oxygen reduction reaction, even though it apparently was below the limit of detection.

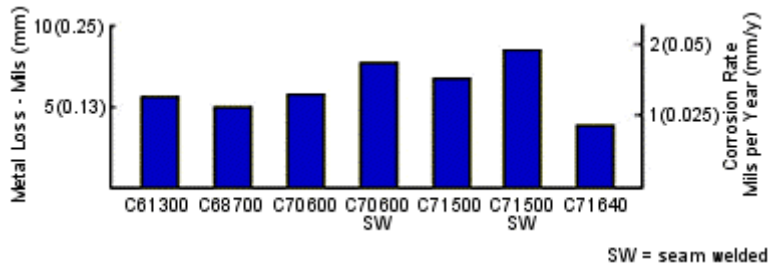
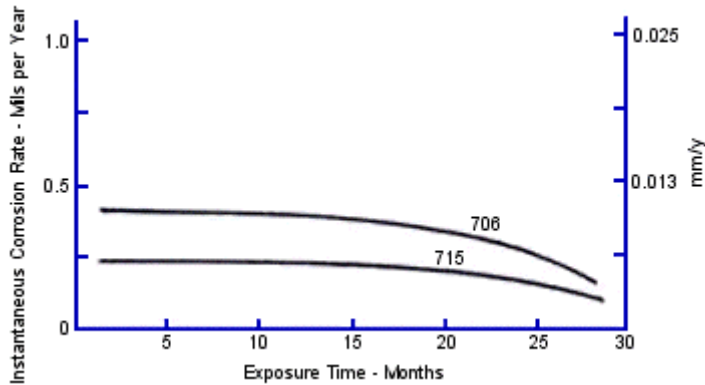


FIGURE 9 - Metal loss (in mils) after 54 months of exposure in the brine heater.(18)



[TDS 35,000ppm; Temperature 220F(104C); Oxygen 0 ppb]

FIGURE 10- Instantaneous corrosion rates (in mils per year) in the recycle brine exchanger.(18)

Corrosion rates in the recycling brine were even lower than in the brine heater, but still not zero as they theoretically should be in the complete absence of oxygen, as shown in Figure 10.

The effect of changes in the amount of oxygen present from 20 to 200 ppb on the corrosion rate of the three copper tube alloys was studied in a special side stream unit, not shown in Figure 6. The corrosion rate of C68700 was found to be more sensitive to changes in low concentrations of oxygen than C70600 and C71500, Figure 11. These data also support the widespread preference for 70600 as compared to C67800 alloy in the heat recovery section of current MSF plants.

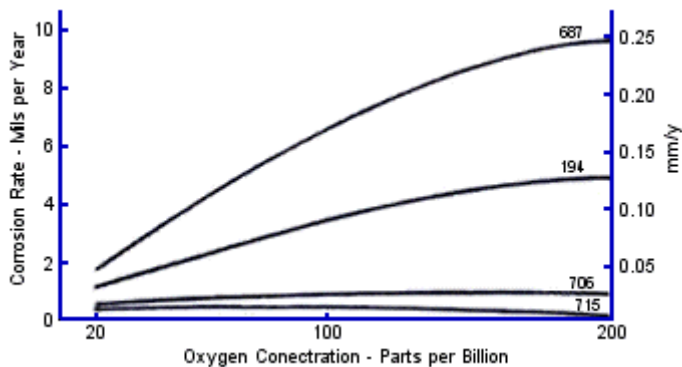


FIGURE 11 - Corrosion rate vs. oxygen concentration after 90 days exposure at 220F (104C). (18)

The feed to the heat recovery unit was unusually low in iron as non-metallic materials were used for the piping and treatment plant. This allowed data to be developed on the effect of iron, as shown in Figure 12. The iron content was adjusted to 0.1 ppm by adding ferrous sulfate. There was a threefold reduction for the three copper alloy tube materials and a reduction of 50% for aluminium bronze in corrosion rate when 0.1 ppm Fe was added to the feed. This beneficial effect of soluble iron on the corrosion resistance of copper alloys in seawater has been reported frequently in the literature by other investigators. It is not surprising that it is also beneficial in the less aggressive heat recovery section environment.

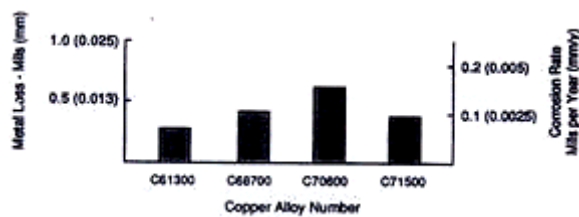


FIGURE 12 - Effect of soluble iron on the metal loss in a Heat Recovery Exchanger after 5 months' exposure (averaged over eight passes). (18)

VAPOR SIDE CORROSION

While the effects of vapour side environments were not studied directly in the Dow Freeport program, the final report does state that unless CO₂ is effectively removed from the feed to the heat recovery section it will form low pH carbonic acid in the vapour zone and lead to increased vapour side corrosion (VSC). Vapour side corrosion of copper nickel tubing in the heat recovery section was not a significant problem in MSP plants until additives displaced acids for treatment.

In additive treated plants, CO₂ is not removed in a decarbonator, as in acid treated plants. In additive treated plants, most of the CO₂ is released in the first several stages downstream of the brine heater. The general approach has been to provide more or larger vents in the first three stages in order to remove the larger volume of CO₂ released. In most cases a simple increase in the venting arrangements has been sufficient to prevent significant vapour side corrosion of copper nickel.

In a few cases, these venting arrangements have proven inadequate. As noted earlier and in the Dow report, should the pH drop low enough due to CO₂, corrosion of copper alloy would be expected provided sufficient air was also present. In the absence of air, CO₂ even at low pH would not corrode copper nickel. In those units where vapour side corrosion has been reported, it appears to be greatest in the lower part of the cold end tube sheet. The film running down the inside of the cold end tube sheet apparently allows enough subcooling for the pH to drop sufficiently for corrosion to occur if there is sufficient air present. As the film runs down the inside of the tube sheet, the driving force to vent the incondensibles decreases

allowing a "bubble of incondensibles" to form and persist in the lower part of the tube bundle at the cold end tube sheet. In those units where this has occurred, it has been found necessary to increase the driving force for removal of noncondensibles by inserting baffles, moving the vent closer to the lower part of the cold end tube sheet, enlarging the vent nearer the cold end tube sheet. The limited number of plants encountering significant VSC indicate that most units have adequate venting systems, or are so well maintained that the inleakage of air is insufficient to support significant vapour side corrosion.

The vent condenser is designed to condense the vapours vented from the steam side of the heat recovery section. Although these vapours are largely steam, they also contain CO₂, oxygen, H₂S if present, and other non-condensibles. Some vent condensers initially were installed with copper alloy tubes as the cooling medium was seawater. Rather severe vapour side corrosion occurred due to the high concentrations of CO₂ and, when present H₂S in the condensing vapours. The use of copper alloys in vent condensers was discontinued in favour of titanium and 6% Mo austenitic stainless steel tubing which are resistant to CO₂ and H₂S.

EXPERIENCE WITH COPPER ALLOYS IN DESALINATION PLANTS

Three detailed surveys of materials performance in desalination plants were made by A. D. Little in 1968, 1972 and 1981.(19,20,21) Polyphosphate and acid were the two treatments in use at the time. These treatments have been largely superseded by the use of anti-scale additives, such as Belgard2000, which permit higher brine heater outlet temperatures and greater efficiency. Despite the change in treatment, the survey data still provides useful information. The three sections of MSF desalination plants are shown in Figure 13 to assist the reader in identifying the different sections in the following tables.

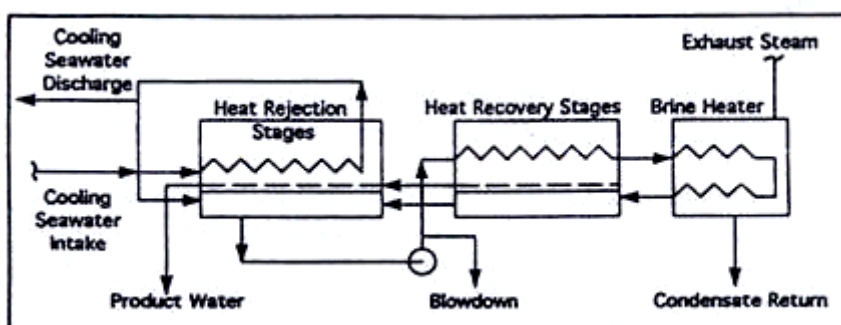


FIGURE 13 - Block flow diagram of a desalting plant.(19)

Two major alterations in the tube corrosion environment occurred with the change to anti-scale treatment. First, the addition of acid during operation and running for long periods at low pHs, as was the practice in a number of acid dosed plants, was discontinued. With introduction of additives acid cleaning at low pH because a separate operation performed for a short period with better control of pH while the plant was off line for cleaning. The change to

anti-scale additives in lieu of acid treatment reduced the metal losses from extended acid cleaning at low pHs, and increased the service life of copper alloy tubing.

Second, there was a change in carbon dioxide release. Carbon dioxide was vented off in the decarbonator in acid treated plants and was not a significant factor in the heat recovery section. In anti-scale additive plants, carbon dioxide remains in the brine until the brine reaches the first flash chamber downstream of the brine heater and may lead to VSC as just discussed.

Data on comparative usage and performance of copper alloy tubing is presented in the following tables. It should be kept in mind that these data include performance in many of the older acid treated plants. As explained above, even better service can be expected in modern anti-scale additive treated plants.

Table 9 reports tubing materials selection for first generation desalination plants. Tube selection for these early plants was based on performance of tubing in coastal power plants and marine condensers.

Table 9. Tube selection in % in existing plants surveyed in 1972.(20)

Tube alloy	Heat Re-ject	Heat Recovery	Brine Heater
C68700 (Al Brass)	23	68	22
C70600 (90-10 CuNi)	34	24.5	49
C71500 (70-30 CuNi)	1.5	1	1.5
C76140 (70-30 CuNi 2Fe-2Mn)	39.5	6	31

By 1976 tube preferences could be based on actual desalination plant experience as well as on projections from coastal power plant and shipboard condenser experience. Tube preferences in 1976 were analysed by recording tube selection preferences in actual bids, as reported in Table 10.(22)

Table 10. Tube usage analysis in % for 75 Mid East desalination plant invitations to bid in 1976, representing some 200 mgpd of capacity. (22)

Tube alloy	Heat Re-ject	Heat Recov-ery	Brine Heater
C68700 (Al Brass)	6	46	0
C70600 (90-10 CuNi)	63	50	67
C76140 (70-30 CuNi 2Fe-2Mn)	27	3	33
Titanium	4	1	0

Materials selection for these 75 plants reflected the early preferences, as reported in Table 9, tempered by some ten years of actual desalination plant operating experience. The principal differences are the dramatic increase in use of C70600 tubing; the equally dramatic decrease

in use of C68700; the major increase in use of C71640 in the heat reject section and the brine heater; and the near phasing out of use of C71500 tubing.

The increase in the use of C71640 in the heat reject section reflects its excellent resistance to erosion-corrosion as well as to sand abrasion in the shallow waters of the Middle East Gulf where sand loading can be high. These properties of C76140 tubing are reflected in the very low failure rates in the heat reject sections as reported in Table 11 below.

The marked increase in the use of C70600 alloy and the decrease in use of C68700 in the large heat recovery section reflects actual experience with the two alloys. Tube failure rates from the 1972 survey by A. D. Little, reported in Table 11 below, show much lower failure rates for C70600 than C68700, especially in plants using acid treatment.

Table 11. Tube failure rates in % for the three sections of 119 MSF Desalination Plants at 80 different locations, from A. D. Littles 1972 survey of materials performance.(20)

Tube alloy	Heat Reject		Heat Recovery		Brine Heater	
	Polyphosphate	Acid	Polyphosphate	Acid	Polyphosphate	Acid
C68700 Al Brass	7.2	17.7	0.18	9.0	5.6	49
C70600 90-10	2.3	2.5	0.6	0.36	0.01	4.2
C76140 70-30 2Fe-2Mn	0.02	0.17	0.025	-	1.57	0

Tube failure rate was calculated by counting the number of tubes plugged as a percent of total tubes. this method of calculating "tube failure rate" overstates the actual tube failure rate by a factor of 5x or more because of the common practice of plugging six to eight "guard" tubes around each leaking tube.

These data indicate much better performance for C70600 than C68700. In the heat recovery section, much of the C70600 tubing was 20 gage, 0.035"(0.9mm), as compared to 18 gage, 0.047"(1.2mm) wall thickness for C68700. Even with a thinner wall, performance of C70600 was significantly better than the heavier-walled C68700, in the plants using acid treatment.

While performance in the heat reject and heat recovery section is believed to reflect alloy capability in these environments, performance in the brine heater may not. Many tube failures in the brine heater are due to mechanical, rather than corrosion factors. When the brine heater scales up, it is necessary to ream out the tubes. The reaming operation not only removes scale but also metal from the wall, sometimes so much metal that the wall is penetrated in the reaming operation.

The higher tube failure rates in the acid treated plants are believed due to two factors: first, removal and reformation of the film with each acid cleaning; and second, unintended extensions of the time at low pHs, when there is no protective film and high corrosion rates.

There have been no major surveys of material performance in plants using high temperature additives. The literature and the author's experience indicate that usage of C70600 tubing has continued to grow relative to C68700. C70600 has performed exceedingly well in the large heat recovery section, except in a few plants where inadequate venting of CO₂ in the first stage has led to high copper in the distillate readings and external corrosion failures of a few tubes. Brine heater tubing has performed well and there has been little change in alloy selection in this stage. Titanium usage has increased in the heat reject section and C70600 usage has decreased. This is believed due to titanium's better resistance to debris, partial blockage, and turbulence, which continue to cause failures of C70600, especially in units where screens are not well maintained. C71640 has held its share and continues to perform well in the heat reject section.

C70600 copper nickel welded piping and solid, clad and sheet lined waterboxes were originally used in the higher temperature heat recovery stages and the brine heater. Usage has expanded to include the full heat recovery section. C70600 tube sheets have been standard throughout the heat recovery section. C71500 tubesheets are used with C71640 tubing. The authors have not encountered problems with copper nickel piping, waterboxes or tubesheets. Performance has been excellent.

Based on 30 years of actual desalination plant experience, and the data developed in industry test programs summarised above, the copper nickel alloys can be expected to give even better service in the future than in the past.

REFERENCES:

1. Lewis, R. O., "the Influence of Biofouling Countermeasures on Corrosion of Heat Exchanger Materials in Seawater", MP, Vol.21, Sept. 1982, pp 3 1-38
2. North, R. F., M. J. Pryor, Corros. Sci., Vol 10, p197, 1970
3. Gilbert, P. T., "A Review of Recent Work on Corrosion Behavior of Copper Alloys in Sea Water", MP Feb., 1982, pp 47-53
4. Pourbaix, M., "Atlas of Electrochemical Equilibria in Aqueous Solutions", NACE, Houston, TX, 1974
5. Tuthill, A. H., "Guidelines for the use of copper alloys in seawater", MP, Vol 26, No. 9pp12-22 1987
6. "Corrosion Resistance of Wrought 90-10 Copper-Nickel-Iron Alloy in Marine Environments", INCO TP A-1222, 1975
7. Ross, W. R., "The effect of seawater composition on corrosion of Cu-Ni-Fe alloys at elevated temperatures", Corrosion/77, Paper No. 94, San Francisco, March 1977.
8. Ross, W. R. and Anderson, D. B., "Hot Sea Water Corrosion of Copper-Base Alloys" MP Sept. 1975, pp 27-32

9. Eford, K. D., "Corrosion/75 Paper No. 54, NACE, Houston, TX 1975.
10. Syrett, B. C., "Sulfide Attack in Steam Surface Condensers", "Proceedings of Second Intl. Conference on Environmental Degradation of Engineering Materials in an Aggressive Environment" 1981 pp 3-14
11. May, T. P. and Weldon, B. A., "Copper-nickel alloys for service in seawater", 24th International Congress on Fouling and Marine Corrosion, Cannes, France, 1964.
12. Thompson, D. H., "Materials Research Std., Vol 1, p 108, 1961
13. Michels, H. T., "The Influence of Corrosion and Fouling on Steam Condenser Performance", J. Materials for Energy Systems, Vol. 1, Dec. 1979
14. Anderson, D. B., "Chlorination of Sea Water-Effects on Fouling and Corrosion", 1. of Engineering for Power, July, 1966
15. Francis, R., "The effect of chlorine additions to cooling water on corrosion of copper alloy condenser tubing," Materials Performance, August, 1982, Vol. 21(8)
16. Sato, S. and Nagata, K., "Factors affecting corrosion of condenser tubes of copper alloys and titanium" Corrosion 77, NACE, Toronto, 1977
17. Hornburg, C. D., "Materials in desalination plants nickel alloys predominate," Desalination Water Reuse Vol. 3/3 pp. 33-40
18. George, P. F., "Copper Alloys in the Desalting Environment" November, 1970 through November, 1973, Dow Chemical Company, Texas Division, Freeport Texas, for the Copper Development Association, 260 Madison Ave. NY, NY 10016
19. Newton, E. H. and Birkett, J. D., "Survey of Materials Behavior in Multi-Stage Flash Distillation Plants", A. D. Little Inc., Cambridge, MA, August, 1968
20. Newton, E. H., Birkett, J. D. and Ketteringham, J. M., A. D. Little Inc., Cambridge, MA, March, 1972
21. Birkett, J. D. and Newton, E. H. "Survey of Service Behavior of Large Evaporative Desalting Plants", A. D. Little Inc., Cambridge, Ma, April 1981
22. Todd, B., Private Communication, 1976